



(12) **EUROPEAN PATENT SPECIFICATION**

(45) Date of publication and mention
of the grant of the patent:
09.04.1997 Bulletin 1997/15

(51) Int Cl.⁶: **C11D 3/395, C11D 3/48**

(21) Application number: **90120218.4**

(22) Date of filing: **22.10.1990**

(54) **Deodorizing and cleaning compositions and method**

Deodorierungs- und Reinigungsmittel sowie Verfahren

Composition et procédé pour la désodorisation et le nettoyage

(84) Designated Contracting States:
DE FR GB IT SE

• **Haber, Mary K.**
Cincinnati, Ohio 45227 (US)

(30) Priority: **23.10.1989 US 425738**

(74) Representative: **Kinzebach, Werner, Dr. et al**
Patentanwälte
Reitstötter, Kinzebach und Partner
Postfach 86 06 49
81633 München (DE)

(43) Date of publication of application:
02.05.1991 Bulletin 1991/18

(73) Proprietor: **S.C. JOHNSON & SON, INC.**
Racine, Wisconsin 53403 (US)

(56) References cited:

GB-A- 2 205 325

US-A- 3 231 503

US-A- 3 466 208

US-A- 3 844 857

US-A- 4 007 262

US-A- 4 790 950

US-A- 4 816 086

(72) Inventors:

• **Hutchings, Richard S.**
Fairfield, Ohio 45014 (US)

EP 0 424 845 B1

Note: Within nine months from the publication of the mention of the grant of the European patent, any person may give notice to the European Patent Office of opposition to the European patent granted. Notice of opposition shall be filed in a written reasoned statement. It shall not be deemed to have been filed until the opposition fee has been paid. (Art. 99(1) European Patent Convention).

Description**FIELD OF THE INVENTION**

5 The present invention refers to aqueous compositions containing an alkali metal halogenite, for example sodium chlorite (NaClO_2) together with a salt of a transition or post transition metal or mixtures thereof and a secondary alcohol. The composition is useful for cleaning and/or odor elimination. It relates also to such compositions containing one or more selected ligands.

BACKGROUND OF THE INVENTION

Aqueous compositions of alkali metal halogenites, especially sodium chlorite, are known. Such compositions appear to function in either of two ways.

15 In one mode of action, the metal chlorite itself acts as an oxidant. For example, when used as an odor eliminator, the chlorite oxidizes the amines, aldehydes and thiols which are the principal components of malodorous mixtures and converts them to odorless oxidation products. In the second mode of action, the halogen dioxide, principally chlorine dioxide, performs the same function.

20 Halogenites such as sodium chlorite when they are a sole component in aqueous solutions at a pH of 10 or above will neutralize odors by oxidation, but only at a low rate. If the pH of the solution is lowered by acidification to a value of about 8 or less, sodium chlorite is converted to chlorine dioxide, a good oxidizer. At the lower pH values the oxidation potential of the sodium chlorite also increases. As the pH decreases further, the rate of formation of chlorine dioxide increases and a series of equilibrium mixtures between chlorine dioxide and the chlorite ion form. The overall oxidizing power of the composition increases because of the greater oxidation potential of the chlorine dioxide and because of the more reactive chlorite species.

25 Unfortunately, chlorine dioxide itself is offending to the olfactory senses even in trace amounts, i.e. about 5 ppm. US-A-4,007,262 discloses odor control compositions having a pH of about 6.5 to 11 and containing an oxidizing agent such as potassium or sodium chlorite together with at least one transition metal salt which is capable of forming an ammonia complex or metal-amine complex.

30 The art, therefore, has been faced with the problem that the reaction kinetics of aqueous compositions of alkali metal halogenites such as sodium chlorite are too slow. On the other hand, useful chlorine dioxide compositions require a low concentration of the chlorine dioxide if they are to avoid the aforementioned drawbacks.

Accordingly, it is an object of this invention to provide aqueous oxidizing compositions comprising an alkali metal halogenite with enhanced reactivity to eliminate odors rapidly but without appreciable formation of halogen oxide gases.

SUMMARY OF THE INVENTION

35 The compositions of this invention are effective for deodorizing any of a variety of malodorous substrates including smoke and other household odors such as toilet and kitchen odors arising from a variety of sources including pets and food wastes or odors from cooking foods, especially burning odors. They are especially useful for deodorizing fabrics such as curtains and upholstery fabrics.

40 Compositions of the present invention comprise an alkali metal halogenite e.g. sodium chlorite in aqueous media together with an effective quantity of at least one salt of a transition or post transition metal, or mixtures of such salts. They also comprise a secondary alcohol to assist in stabilizing the sodium chlorite.

45 The compositions may be used, for example, by spraying, or otherwise contacting, a fabric or other substrate from a container having a spray-pump mechanism, to effect deodorization. They also may be utilized as cleaning compositions.

The pH of the composition is largely dependent upon the constituents used. Typically, the pH will be 4 to 11. Preferably, so as to avoid or limit the formation of chlorine dioxide, the pH should be slightly acidic or basic, most preferably above about 8.

50 The active components of the compositions of this invention function principally by reacting with the odor causing chemicals in the area to be treated to form reaction products that are odorless or at least non-offending. They thus reduce odors by chemical reaction, not by masking the odors with perfumes. However, perfumes may be included in the compositions to make them more pleasant to work with or, in some cases, to mask certain odors which are not neutralized by oxidation.

55 The alkali metal halogenite typically is present in an amount of at least 0.01%, sodium chlorite being the preferred oxidizing agent. Typically, the halogenite is present in an amount of 10% or less, although this is not intended to limit the upper concentration level.

The compositions will normally contain a sufficient level of the selected transition or post transition salt to effect

an increase in the oxidation potential of the sodium chlorite. Wide variations are possible, depending on the salt selected. Of course, in some instances the salt chosen may also provide a second function, in which event the level for that function will govern. Typically, the metal ion will be present in the aqueous solution at a level of at least 0.0001% by weight. Amounts in excess of about 1% will not usually provide any additional benefit.

The perfumes, when employed in the compositions of the invention should be stable, i.e. not oxidized by the halogenite or, if unstable, should be employed together with a stabilizing amount of an anionic surfactant as described in copending and commonly owned patent application, serial number 54338 filed May 26, 1987, corresponding to GB-A-2 205 325.

A perfume, if employed, is normally utilized at a low concentration, i.e. 0.25% or less. If a perfume which is unstable in aqueous halogenite compositions is utilized, it will be in association with a stabilizing amount of an anionic surfactant, typically 10% or less of the selected surfactant.

In this disclosure and the appended claims, all amounts of the components of the compositions are defined in weight percent based on the total weight of the composition.

DETAILED DESCRIPTION OF THE INVENTION AND PREFERRED EMBODIMENTS

Attempts to prepare deodorant compositions containing alkali metal halogenites have been hampered for a number of reasons including those discussed above. In addition, the alkali metal halogenites, at least when compared to chlorine dioxide, are not considered to be strong oxidizers.

It has now been discovered that such compositions can be prepared in a useful form by including selected quantities of salts of transition and post transition metals or mixtures thereof. It appears that the transition or post transition metal ion may in some manner enhance the oxidation potential of the chlorite ion species, but this is not necessarily the mode of action of these salts.

Accordingly, the compositions of this invention are aqueous oxidizer compositions typically having a pH of from 4 to 11, containing 0.01% to 10% preferably from 0.05% to 5% of an alkali metal halogenite together with at least 0.0001% of an ion of a transition or post transition metal or mixtures thereof and a secondary alcohol. The compositions may optionally contain a perfume which may be a reactive perfume stabilized by an anionic surfactant.

In some instances the formation of chlorine dioxide is permitted, such compositions also containing the alkali metal halogenite and the transition metal or post transition metal salt ion. In such compositions the pH is in the lower region of the pH range, i.e. from 4 to 8. The presence of the chlorine dioxide does not appear to negatively influence the enhancement of oxidation potential enjoyed by the alkali metal halogenite and provided by the presence in solution of the salt ions.

It has been observed that the basic compositions of this invention can be improved utilizing certain ligands and secondary alcohols. The alcohols can be employed alone or together with the ligand. Both of these additives stabilize the compositions by inhibiting the decomposition of the chlorite, the formation of chlorine dioxide and the formation of precipitates.

The presently preferred secondary alcohols are isopropanol and secondary butyl alcohol. When employed, with or without a ligand the concentration of the alcohol is normally up to about 15% preferably 1% to 15% although appreciable variations are possible.

The term ligand as used in this disclosure and claims means a molecule, ion or atom that interacts with the transition or post transition metal salt and inhibits the formation of chlorine dioxide even at pH levels where chlorine dioxide would normally be produced. Typically ligands as employed herein are compounds which will form complexes with the central atom. Suitable ligands include oxalic, citric, acetic, polyacrylic and ethylenediamine tetraacetic acids, the alkali metal salts thereof and ammonia. Suitable concentrations of the ligand are about 10 times that of the metal salt. Therefore, the concentration range is from 0.001% to 10%.

In the presently preferred practice of this invention for use as deodorizers, aqueous sodium chlorite compositions are employed which do not produce in excess of trace amounts of chlorine dioxide within one month of preparation, preferably three months, most preferably six months. By trace concentration is meant a level of chlorine dioxide detectable by ordinary analytical methods and/or by olfactory sensory evaluation, usually less than about 5 ppm, preferably less than 1 ppm, in said composition.

Sodium chlorite is preferably present in an amount of from 0.01 to 10%, most preferably 0.05 to 1%. Compositions containing more than 10% are not appreciably more efficient than those containing less than this amount, but are within the scope of this invention.

Transition and post transition metals are a defined class of metals including metals that have filled or partially filled outer d orbitals, while post transition metals are those that have filled d orbitals and filled or partially filled p orbitals. Any of a variety of inorganic or organic salts of these metals may be usefully employed. Inorganic salts include, for example, halogen salts such as bromide or chloride salts, sulfate salts, phosphate salts and nitrate salts or salts with any other suitable anion. Organic salts include, for example, citrates, oxalates, acetates, and carboxylates such as

stearate. Solubility of these salts is not a major factor so long as they ionize to some extent, to provide the requisite concentration of ions in solution. Of course, undissolved salt may be unattractive to the consumer, or may interfere with the dispensing mechanism of the product dispenser. The presently preferred salts are chloride salts such as cupric, zinc, stannic, or ferric chloride; sulfate salts such as cupric sulfate or citrate salts such as cupric citrate.

The transition and post transition metal salts are preferably present in the compositions in an amount suitable to provide a metal ion concentration of above 0.0001%. Typically the metal ion concentration is from 0.01 to 1%, preferably from 0.05 to 0.1%. As with the chlorite concentration, there is no appreciable advantage in employing more than 1% of the selected salts. Therefore, it is preferred not to do so unless a particular salt also has a second functionality.

Certain transition or post transition metal salts will acidify the chlorite compositions of the present invention. For example, zinc chloride appears to ionize such that the zinc ion forms a coordination sphere with hydroxyl ions from the water constituent. Consequently, there is an excess of H^+ ions and a lowering of pH.

When the pH falls to less than about 8, chlorine dioxide begins to form, as previously noted. Because the conversion of chlorite to chlorine dioxide is pH dependent, the chlorine dioxide concentration may be regulated by regulating pH, for instance, with buffers as hereinafter explained.

Generally, the presence of chlorine dioxide is not preferred. When present in the compositions, however, it is at a level of less than about 5ppm, preferably less than about 1ppm.

Buffers are used to maintain pH within certain limits. In one aspect of the invention, the buffer may be selected to maintain pH above about 8. In such instances chlorine dioxide is not formed in detectable amounts. In another aspect, the pH level maintained by the buffer is such that a useful, but non-toxic concentration of chlorine dioxide is achieved.

Typically, the buffer or other pH adjustor will be present in an amount of less than about 10% by weight of the composition. Preferably, when utilized, the buffer will be from about 0.1 to about 5% of the composition, and will be in a sufficient quantity to maintain a basic pH. Suitable buffers include sodium carbonate, sodium borate, sodium bicarbonate and calcium hydroxide.

Alternatively, the solution may be initially prepared as a strongly alkaline solution with a strong alkali such as an alkali metal hydroxide, suitably sodium or potassium hydroxide.

If a perfume is employed in the practice of this invention, it should be stable or stabilizable in the presence of the alkali metal halogenite. Most perfumes which are useful in compositions of this nature contain oxidizable groups so that, as a practical matter, the selected perfume will be one which is stabilized by the use of an anionic surfactant as described in the aforementioned U.S.S.N. 54338 corresponding to GB-A-2 205 325, the disclosure of which is incorporated herein by reference.

Optionally, other adjuvants may be included in the compositions of the present invention, provided that such adjuvants do not exhibit incompatibility. For example, surfactants, colorants, chelating agents, sequestering agents, builders, and the like may be included. Care, however, must be exercised when selecting these optional materials to ensure compatibility. Thus, adjuvants with aldehyde or other easily oxidizable groups should be avoided since they will be oxidized by the halite present and undesirably decrease its concentration. Aldehydes will be oxidized to acids which will also increase the production of the halogen dioxide for the reasons explained above. This is normally undesirable. Stabilization of compositions containing colorant is disclosed in application U.S.S.N. 54347 corresponding to GB-A-2 205 326. The anion surfactants disclosed in GB-A-2 205 325 and GB-A-2 205 326 are especially preferred. The above mentioned adjuvants are incorporated in an amount suitable for attaining their particular function, typically in an amount of less than about 1%, except for surfactants which are generally less than about 10% preferably less than about 5%.

The compositions of this invention are useful for the elimination of a variety of odors such as pet odors and trash can odors. They are generally useful as room deodorants particularly for the elimination of smoke odors from soft surfaces including fabrics such as the fabrics in upholstery, carpets, curtains and drapes. Preferably, they are dispensed from a spray dispenser which provides a fine mist of the product. These compositions may also be used as general purpose cleaning/disinfectant agents for laundry and other household and janitorial cleaning purposes.

The following examples are given by way of illustration only and are not to be regarded as limitations of this invention many apparent variations of which are possible without departing from the spirit and scope thereof.

Except where noted, in the examples, the sodium chlorite was 80% active, the remainder being 5% sodium hydroxide and 15% sodium chloride.

EXAMPLE 1 (Reference example)

The following compositions were made in accordance with the present invention. Each of these formulations contained 0.007 mols of the metal ion indicated, 0.5% sodium chlorite and water q.s. 100%.

Metal	pH
$FeCl_3 \cdot 6H_2O$	4.0

(continued)

Metal	pH
CrCl ₃ · 6H ₂ O	4.1
Zn Cl ₂	6.2
Cu ₂ (C ₆ H ₄ O ₇)	9.5
CuSO ₄ · 5H ₂ O	5.2
CuCl ₂ · x H ₂ O	4.5

EXAMPLE 2 (Reference example)

This example illustrates the stabilization of the metal salt containing composition of the invention with the ligand sodium citrate.

Control Formula		
Sodium Chlorite	0.4%	
Metal salt	0.2%	
Deionized water q.s		
Metal Salt	Control pH No ligand	pH with 2.3% Added Sodium Citrate
Sn Cl ₂ · 2H ₂ O	3	6
Cr Cl ₃ · 6H ₂ O	4	6
Cu Cl ₂ · XH ₂ O	6	4
Fe Cl ₃ · 6H ₂ O	4	7
Zn Cl ₂	6	8
CuSO ₄ · 5H ₂ O	6	7

Chlorine dioxide was produced in all of the controls. None of the compositions containing citrate evidenced generation of chlorine dioxide even though the copper chloride, tin chloride and chromium chloride composition all were at an acid pH, and the other compositions were either neutral or slightly basic.

EXAMPLE 3 (Reference Example)

This example is similar to Example 2 except that ethylenediamine tetraacetate (EDTA) was used as a ligand. No precipitates were formed in any of the EDTA containing solutions and no chlorine dioxide was produced.

Control Formula		
Sodium Chlorite	0.4%	
Metal salt	0.2%	
Deionized water q.s		
Metal Salt	Control pH No ligand	pH with 2.3% Added EDTA
Sn Cl ₂ · 2H ₂ O	3	9
Cr Cl ₃ · 6H ₂ O	4	9
Fe Cl ₃ · 6H ₂ O	4	11
Zn Cl ₂	6	11
Cu Cl ₂ · XH ₂ O	6	11

EXAMPLE 4 (Reference Example)

This example is similar to Examples 2 and 3 and shows the use of sodium oxalate as a ligand stabilizer. No precipitates or chlorine dioxide formed in any of the oxalate containing compositions.

Control Formula		
Sodium Chlorite	0.4%	
Metal salt	0.2%	
Deionized water q.s		
Metal Salt	Control pH No ligand	pH with 0.2% Added Sodium Oxalate
Sn Cl ₂ . 2H ₂ O	3	6
Fe Cl ₃ . 6H ₂ O	4	8
Cu So ₄ . 5H ₂ O	5	10
Zn Cl ₂	6	9
Cu Cl ₂ . XH ₂ O	5	10

EXAMPLE 5

The following compositions were prepared to illustrate compositions of the invention containing a metal salt, a ligand and a perfume. Composition C additionally contains a surfactant.

A	B	C
82.07% deionized water	82.03% deionized water	80.49% deionized water
0.5% sodium chlorite	0.5% sodium chlorite	0.5% sodium chlorite
2.32% sodium citrate	2.32% sodium citrate	2.32% sodium citrate
0.1% zinc chloride	0.1% zinc chloride	0.1% zinc chloride
15% isopropyl alcohol	15% isopropyl alcohol	15% isopropyl alcohol
0.01% perfume	0.05% perfume	0.05% perfume

EXAMPLE 6

This example shows the efficacy of certain ligands to stabilize the basic compositions of this invention both with respect to pH and retention of sodium chlorite in the composition.

Base Formula	Percent Ligand			
Sodium Chlorite 0.4%	Citric Acid 0.75%			
Isopropyl Alcohol 5%	Sodium Citrate 2.32%			
Zinc Chloride 0.1%	Sodium Acetate 0.33%			
Ligand	EDTA 1.54%			
Deionized Water Q.S.	Sodium Polyacrylate 0.02% (PAA)			
Ligand	Initial PH	Final PH 28 DAYS/125° F	Final PH 100 F/ 90 DAYS	Final PH Room Temp 90 DAYS
Citric Acid	7.2	7.2	9.1	9.2
Citrate	7.6	7.8	8.1	8.1
Acetate	6.8	6.2	6.4	6.5
EDTA	11.1	9.7	9.8	9.3
PAA	6.5	6.0	6.6	6.5

EP 0 424 845 B1

Ligand	PPM Chlorite Initial	PPM Chlorite 28 DAYS 125 F	PPM Chlorite 100 F/90 DAYS	PPM Chlorite Room Temp
Citric Acid	3300	65	460	3000
Citrate	3200	3100	3100	3200
Acetate	3400	2600	3000	3600
EDTA	3300	2800	3200	3000
PAA	3300	2300	2800	3200
1. No chlorine dioxide was formed 2. All room temperature samples stable to chlorite loss 3. The acetate and PAA sample maintained acidic pH's				

EXAMPLE 7

Cloths 1 to 11 below were treated with the recited compositions and rated for efficacy of smoke odor removal.

1. 0.25% sodium chlorite, 0.05% copper citrate
2. 0.1% sodium chlorite, 0.01% copper citrate
3. 0.5% sodium chlorite, 0.1% chromium chloride, 0.1% sodium hydroxide
4. 0.5% sodium chlorite, 0.1% tin chloride, 0.1% sodium hydroxide
5. 0.5% sodium chlorite, 0.1% cobalt chloride
6. No smoke odor applied
7. Unidentified smoked cloth
8. Water
9. 2.0% sodium chlorite, 1.0% copper citrate
10. 0.5% sodium chlorite, 0.1% copper citrate
11. 0.5% sodium chlorite

The efficacy of these compositions was evaluated as follows. Test cheese cloths were saturated with smoke odor in a test chamber for about one hour. Equal amounts of the above compositions were applied by spraying to the test cloths. Cloths 6 and 7 were used as controls. The smoke treated cloths were rated for smoke odor blindly by a panel of judges, the cloth's acting as a basis of comparison. A suitable number of observations were made so that the results were statistically significant to a 95% confidence level. The values above are reported on a scale normalized to 0 to 10 (zero meaning total perceived smoke odor elimination).

The results are shown below.

Cloth	Odor Rating 30 minutes	Odor Rating 3 hours	Retained Smoke Odor 3 hours
1	8.2	1.9	yes
2	7.9	3.3	No
3	3.9	0.9	No
4	3.7	0.9	No
5	9.0	3.1	No
6	2.4	0	No
7	4.7	5.3	Yes
8	9.8	0.7	Yes
9	2.9	1.1	No
10	2.4	0.7	No
11	7.2	2.2	Yes

EXAMPLE 8

The following compositions were evaluated for reduction of smoke odor using the same procedure as in Example 7 except that the samples were allowed to air dry for 5 hours after treatment.

Composition	Odor Rating
5% sodium chlorite plus 1% copper citrate	0.64
5% sodium chlorite plus 0.1% copper citrate	0.86
0.5% sodium chlorite plus 0.1% zinc chloride	1.57

EXAMPLE 9

This example illustrates the ability of certain alcohols to stabilize the sodium chlorite in the basic compositions of this invention.

Control Formula

0.4%	sodium chlorite (100% active basis)
0.1%	zinc chloride
99.5%	deionized water

Test Formulas

5%	alcohols (Samples B-K)
0.4%	sodium chlorite (100% active basis)
0.1%	zinc chloride
94.5%	deionized water

14 Day Time Period		
Formula #/Alcohol	System pH	Chlorine Dioxide Present
A. Control	6.4	Yes
B. Isopropanol	6.4	No
C. Ethanol	6.5	Yes
D. Methanol	6.4	Yes
E. N-Propanol	6.5	Yes
F. N-Butanol	6.6	Yes
G. Isobutanol	6.6	Yes
H. 2-Butanol	6.2	No
I. Tert-butanol	6.5	Yes
J. Ethylene glycol	6.9	Yes
K. Propylene glycol	6.7	Yes

Note that no ligands are present in these formulations. As will be seen from the above only secondary alcohols stabilized the compositions as evidenced by the fact that chlorine dioxide was released in all other compositions as would normally be expected since all compositions were acidic.

This example was repeated with similar compositions utilized ethanol, isopropyl alcohol and secondary butyl alcohol and tap water instead of deionized water since it was postulated that the dissolved chlorine in the tap water was responsible for the destabilization of the sodium chlorite. Again it was observed that the secondary alcohols were useful as stabilizing agents.

Claims

1. An aqueous oxidizer composition having a pH of from 4 to 11 comprising on a weight basis

- a) from 0.01 to 10% of an alkali metal halogenite;
- b) a salt selected from the group consisting of salts of transition and post transition metals and mixtures thereof, said salt being present in a concentration sufficient to obtain a metal ion concentration of at least 0.0001% and
- c) a secondary alcohol.

5

2. The composition of claim 1 wherein the alkali metal halogenite is sodium chlorite.
3. The composition of claim 2 wherein the salt cation is selected from the group consisting of zinc, stannic, ferric, cupric salts and wherein the salt anion is selected from the group consisting of chloride, sulfate, citrate and phosphate.
4. The composition of claim 2 wherein the sodium chlorite is present in an amount of from 0.05 to 1%.
5. The composition of claim 2 or 4 wherein the pH is above 8.
6. The composition of claim 5 including a buffer effective to maintain the pH above 8.
7. The composition of claim 6 wherein the buffer is selected from the group consisting of sodium carbonate, sodium borate, sodium bicarbonate and calcium hydroxide.
8. The composition of claim 5 wherein the salt is selected from the group consisting of zinc chloride, stannic chloride, ferric chloride, cupric chloride, cupric sulfate and cupric citrate.
9. The composition of claim 5 wherein the salt is cupric citrate.
10. The composition of anyone of the preceding claims additionally containing from 0.001% to 10% of a ligand.
11. The composition of anyone of the preceding claims wherein the secondary alcohol is present in an amount of 1 to 15%.
12. The composition of anyone of the preceding claims wherein the secondary alcohol is isopropanol or secondary butyl alcohol.
13. The composition of claim 10 wherein the ligands are selected from the group consisting of oxalic, citric, acetic, polyacrylic and ethylenediamine, tetracetic acids and alkali metal salts thereof and ammonia.
14. A method of deodorizing malodorous substrates which comprise contacting the substrate with an amount of a composition of anyone of claims 1 to 13 which is sufficient to effect deodorization.
15. A method of cleaning/disinfecting a substrate which comprises contacting the substrate with a cleaning/disinfecting amount of a composition of anyone of claims 1 to 13.

10

15

20

25

30

35

40

Patentansprüche

45

1. Wäßrige Oxidationsmittelzusammensetzung mit einem pH von 4 bis 11, welche auf Gewichtsbasis umfaßt:
 - a) 0,01 bis 10 % eines Alkalimetallhalogenits,
 - b) ein Salz, das ausgewählt ist unter den Salzen von Übergangsmetallen und Post-Übergangsmetallen und Gemischen davon, wobei das Salz in einer Konzentration vorhanden ist, die ausreicht, um eine Metallionenkonzentration von wenigstens 0,0001 % zu erhalten, und
 - c) einen sekundären Alkohol.
2. Zusammensetzung nach Anspruch 1, worin das Alkalimetallhalogenit Natriumchlorit ist.
3. Zusammensetzung nach Anspruch 2, worin das Salzkation ausgewählt ist unter Zink-, Zinn-IV-, Eisen-III- und Kupfer-II-Salzen und worin das Salzanion ausgewählt ist unter Chlorid, Sulfat, Citrat und Phosphat.

55

4. Zusammensetzung nach Anspruch 2, worin das Natriumchlorit in einer Menge von 0,05 bis 1 % vorhanden ist.
5. Zusammensetzung nach Anspruch 2 oder 4, worin der pH-Wert höher als 8 ist.
- 5 6. Zusammensetzung nach Anspruch 5, enthaltend einen Puffer, durch welchen der pH-Wert höher als 8 gehalten wird.
7. Zusammensetzung nach Anspruch 6, worin der Puffer ausgewählt ist unter Natriumcarbonat, Natriumborat, Natriumbicarbonat und Calciumhydroxid.
- 10 8. Zusammensetzung nach Anspruch 5, worin das Salz ausgewählt ist unter Zinkchlorid, Zinn-IV-chlorid, Eisen-III-chlorid, Kupfer-II-chlorid, Kupfer-II-sulfat und Kupfer-II-citrat.
9. Zusammensetzung nach Anspruch 5, worin das Salz Kupfer-II-citrat ist.
- 15 10. Zusammensetzung nach einem der vorhergehenden Ansprüche, zusätzlich enthaltend einen Liganden in einer Menge von 0,001 % bis 10 %.
11. Zusammensetzung nach einem der vorhergehenden Ansprüche, worin der sekundäre Alkohol in einer Menge von 1 bis 15 % vorhanden ist.
- 20 12. Zusammensetzung nach einem der vorhergehenden Ansprüche, worin der sekundäre Alkohol Isopropanol oder sekundärer Butylalkohol ist.
- 25 13. Zusammensetzung nach Anspruch 10, worin die Liganden ausgewählt sind unter Oxal-, Citronen-, Essig-, Polyacryl- und Ethylendiamintetraessigsäure und den Alkalimetallsalzen davon und Ammoniak.
14. Verfahren zum Desodorieren übelriechender Substrate, wobei man das Substrat mit einer zur Erzielung einer Desodorierung ausreichenden Menge einer Zusammensetzung nach einem der Ansprüche 1 bis 13 in Kontakt bringt.
- 30 15. Verfahren zum Reinigen/Desinfizieren eines Substrates, wobei man das Substrat mit einer reinigenden/desinfizierenden Menge einer Zusammensetzung nach einem der Ansprüche 1 bis 13 in Kontakt bringt.

Revendications

1. Composition aqueuse oxydante ayant un pH compris entre 4 et 11, comprenant, par rapport à une base massique :
 - 40 a) de 0,01 à 10 % d'un halogénite de métal alcalin ;
 - b) un sel choisi dans le groupe constitué des sels des métaux de transition et de post-transition et de leurs mélanges, ledit sel étant présent en une concentration suffisante pour obtenir une concentration en ion métallique d'au moins 0,0001 % ; et
 - 45 c) un alcool secondaire.
2. Composition selon la revendication 1, dans laquelle l'halogénite de métal alcalin est le chlorite de sodium.
3. Composition selon la revendication 2, dans laquelle le cation du sel est choisi dans le groupe constitué des sels de zinc, stanniques, ferriques, cuivriques et dans laquelle l'anion du sel est choisi dans le groupe constitué du chlorure, du sulfate, du citrate et du phosphate.
- 50 4. Composition selon la revendication 2, dans laquelle le chlorite de sodium est présent en une quantité de 0,05 à 1 %.
5. Composition selon la revendication 2 ou 4, dans laquelle le pH est supérieur à 8.
- 55 6. Composition selon la revendication 5, comprenant un tampon efficace pour maintenir le pH à plus de 8.
7. Composition selon la revendication 6, dans laquelle le tampon est choisi dans le groupe constitué du carbonate

de sodium, du borate de sodium, du bicarbonate de sodium et de l'hydroxyde de calcium.

- 5
8. Composition selon la revendication 5, dans laquelle le sel est choisi dans le groupe constitué du chlorure de zinc, du chlorure stannique, du chlorure ferrique, du chlorure cuivrique, du sulfate cuivrique et du citrate cuivrique.
9. Composition selon la revendication 5, dans laquelle le sel est le citrate cuivrique.
- 10
10. Composition selon l'une quelconque des revendications précédentes, contenant de plus de 0,001 à 10 % d'un ligand.
11. Composition selon l'une quelconque des revendications précédentes, dans laquelle l'alcool secondaire est présent en une quantité de 1 à 15 %.
- 15
12. Composition selon l'une quelconque des revendications précédentes, dans laquelle l'alcool secondaire est l'isopropanol ou l'alcool butylique secondaire.
- 20
13. Composition selon la revendication 10, dans laquelle les ligands sont choisis dans le groupe constitué des acides oxaliques citrique, acétique, polyacrylique et éthylènediaminetétracétique et leurs sels de métaux alcalins et l'ammoniac.
- 25
14. Procédé de désodorisation de substrats malodorants qui comprend la mise en contact du substrat avec une quantité d'une composition de l'une quelconque des revendications 1 à 13 qui est suffisante pour réaliser la désodorisation.
- 30
15. Procédé de nettoyage/désinfection d'un substrat qui comprend la mise en contact du substrat avec une quantité nettoyante/désinfectante d'une composition de l'une quelconque des revendications 1 à 13.
- 35
- 40
- 45
- 50
- 55